

# Modern Chemistry Section 3 Gases Answer Key

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### **Modern Chemistry Section 3 Gases**

Example 1.3; A 3.80 g of oxygen gas in a pump has volume of 150 mL. constant temperature and pressure. ... Since the attractive forces between molecules do exist in real gases, the pressure of real gases is actually lower than of the ideal gas equation. ... Petrucci, Ralph H. General Chemistry: Principles and Modern Applications. 9th Ed. Upper ...

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## **Gas Laws: Overview - Chemistry LibreTexts**

Nuclear chemistry is the sub-field of chemistry dealing with radioactivity, nuclear processes, and transformations in the nuclei of atoms, such as nuclear transmutation and nuclear properties.. It is the chemistry of radioactive elements such as the actinides, radium and radon together with the chemistry associated with equipment (such as nuclear reactors) which are designed to perform nuclear ...

## **Nuclear chemistry - Wikipedia**

The BYU'S chemistry section hosts over 1500 chemistry articles for students to use as free study resources. Links to each of these articles have been sorted under their parent concepts and can be found in the collapsible tables provided below.

## **Chemistry - Introduction, Branches, Concepts, Free Resources**

Naming Ions. The name of a monatomic

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cation is simply the name of the element followed by the word ion. Thus,  $\text{Na}^+$  is the sodium ion,  $\text{Al}^{3+}$  is the aluminum ion,  $\text{Ca}^{2+}$  is the calcium ion, and so forth.. We have seen that some elements lose different numbers of electrons, producing ions of different charges (Figure 3.3).

### **5.7: Naming Ionic Compounds - Chemistry LibreTexts**

Let us consider one mole of a monoatomic gas in thermal equilibrium at temperature  $T$ . Each molecule has 3 degrees of freedom due to translatory motion. According to kinetic theory of gases, the mean kinetic energy of a molecule is  $\frac{3}{2} kT$ .  $\frac{1}{2} mC_x^2 = \frac{1}{2} mC_x^2 + \frac{1}{2} mC_y^2 + \frac{1}{2} mC_z^2$ . So,  $\frac{1}{2} mC_x^2 + \frac{1}{2} mC_y^2 + \frac{1}{2} mC_z^2 = \frac{3}{2} kT$

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